noncoordinating ether function occurs at higher wavenumbers; the less intense band is caused by a chelating ether phosphine.

Complex **2** does not react with H_2 at atmospheric pressure,[12] the proton NMR spectrum of the complex dissolved in acetone remaining virtually unchanged after 20 min of hydrogen flushing. However, activation of the ruthenium-oxygen bond proceeds with addition of another 1 equiv of $Ph_2PCH_2CH_2OCH_3$, yielding the hydrido compound *cis*-$H_2Ru(P\sim O)_4$ (**3**). The cis configuration is obvious in the $^{31}P\{^1H\}$ NMR spectrum: two identical triplets for pairs of cis and trans phosphorus nuclei appear.

In conclusion, **2** represents an excellent model complex to demonstrate the opening and closing mechanism of the P,O ligands under very mild conditions.

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(12) Knoth, W. H. *J. Am. Chem. Soc.* **1972**, *94*, 104.

Registry No. **1**, 109011-62-3; **2**, 111557-09-6; **3**, 111557-10-9.

Additions and Corrections

1987, Volume 26

Laurence D. Rosenhein and John W. McDonald*: Synthesis and Characterization of the $[(CO)_4MoS_2MS_2]^{2-}$ and $[(CO)_4MoS_2MS_2Mo(CO)_4]^{2-}$ Ions (M = Mo, W): Species Containing Group VI (6) Metals in Widely Separated Formal Oxidation States.

Page 3414. At the end of the introduction section, we should have quoted the following reference as a prior example of a dinuclear complex containing group 6 metals in widely separated oxidation states: Cotton, F. A.; Schwotzer, W. *J. Am. Chem. Soc.* **1983**, *105*, 5639.—John W. McDonald